

## Mass-Mol / Mol - Mass Stoichiometry

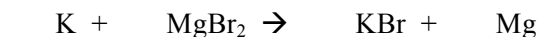
Write out the flowcharts for each type of stoichiometry that you have in your notes in the space below.

### Mass-Mol Stoichiometry

### Mol – Mass Stoichiometry

Balance the following reactions and then use them to answer the questions below.

**Show all your work with units.**



If 2.0 **grams** of K are used in the above reaction, how many **mol** of Mg are produced?

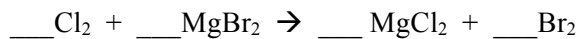
1. What type of stoichiometry is this? \_\_\_\_\_
2. The first step is to convert grams of K to mol of K, do that below.
3. Now use a molar ratio from the balanced equation to convert from mol of K to mol of Mg.

Follow the same procedures as 1-3 and use the flow charts above to answer the following questions:

4. In the above reaction, how many mol of KBr are produced when 13 g of K are used?
5. If 46.0 grams of  $\text{MgBr}_2$  are reacted, how many mol of Mg are produced?

Balance the following reactions and then use them to answer the questions below.

**Show all your work with units.**



If 2.0 **mol** of  $\text{Cl}_2$  are used in the above reaction, how many **grams** of  $\text{Br}_2$  are produced?

1. What type of stoichiometry is this? \_\_\_\_\_
2. The first step is to use a molar ratio to convert mol of  $\text{Cl}_2$  to mol of  $\text{Br}_2$ , do that below.
3. Now use molar mass to convert mol of  $\text{Br}_2$  to grams of  $\text{Br}_2$ .

Follow the same procedures as 1-3 and use the flow charts from the other side to answer the following questions:

4. In the above reaction, how many grams of  $\text{MgCl}_2$  are produced when 0.50 mol of  $\text{Cl}_2$  are used?
5. If 5.17 mol of  $\text{MgBr}_2$  are reacted, how many g of  $\text{Br}_2$  are produced?